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 London Dispersion Forces. The London dispersion force is the weakest intermolecular force. The London dispersion force is a temporary attractive force that results when the electrons in two adjacent atoms occupy positions that make the atoms form temporary dipoles.

[London dispersion force - Wikipedia](#)
 London dispersion forces (LDF, also known as dispersion forces, London forces, instantaneous dipole induced dipole forces, or loosely van der Waals forces) are a type of force acting between atoms and molecules.

[Dispersion Force | Introduction to Chemistry](#)
 These London dispersion forces are often found in the halogens (e.g., F₂ and I₂), the noble gases (e.g., Ne and Ar), and in other non-polar molecules, such as carbon dioxide and methane. London dispersion forces are part of the van der Waals forces, or weak intermolecular attractions.

[Dispersion Forces - definition of Dispersion Forces by The ...](#)
 Dispersion Forces synonyms, Dispersion Forces pronunciation, Dispersion Forces translation, English dictionary definition of Dispersion Forces. n. A weak, short-range attractive force between atoms or molecules caused by their dipole moments, often arising in otherwise nonpolar atoms or molecules.

[London Dispersion Force Definition - ThoughtCo](#)
 The London dispersion force is the weakest of the van der Waals forces and is the force that causes nonpolar atoms or molecules to condense into liquids or solids as temperature is lowered. Even though it is weak, of the three van der Waals forces (orientation, induction, dispersion), the dispersion forces are usually dominant.

[Definition Of Dispersion Forces | Free Physics dictionary ...](#)
 London dispersion force is a weak intermolecular force between two atoms or molecules in close proximity of each other More About Dispersion Forces Atoms and nonpolar molecules are electrically symmetrical and have no dipole moment because their electronic charge cloud is symmetrically distributed.

[10.1 Intermolecular Forces Chemistry - opentextbc.ca](#)
 All of these compounds are nonpolar and only have London dispersion forces; the larger the molecule, the larger the dispersion forces and the higher the boiling point. The ordering from lowest to highest boiling point is therefore C₂H₆ < C₃H₈ < C₄H₁₀.

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[Teaching Students Who Are Exceptional Diverse And At Risk In The General Education Classroom Fifth Edition Victorian Popular Fiction 1860 80 Microsoft Office Access 2007 Complete Concepts And Techniques 1st Edition](#) 016 - London Dispersion Forces In this video Paul Andersen describes the positive force intermolecular forces found between all atoms and molecules. As electrons are distributed unevenly it